# MIXTURES AND MATTER TEST REVIEW NOTES

| VOCABULARY TERM OR CONCEPT | WHAT YOU NEED TO KNOW |
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| Mixture | Made of two or more substances – elements, compounds, or both together at the same place but not chemically combined. |
| Heterogeneous Mixture | A mixture in which you can see the different parts.<br>• Be able to indentify some common heterogeneous mixtures. (salad, soil, etc.) |
| Suspension | An example of a heterogeneous mixture in which a solid will settle out of a liquid over time, sand and water mixed together creates a suspension. |
| Homogeneous Mixture | A mixture in which the substances are mixed so evenly that you cannot see the different parts of the mixture. (water & food coloring, Saltwater) |
| Colloid | A homogeneous mixture in which one substance stays suspended in another substance but does not dissolve or settle to out. Milk is an example of a colloid mixture. |
| Solution | A homogeneous mixture which contains a solute and a solvent. The solvent dissolves the solute(s) creating a mixture which appears to be all the same substance. |
| | • Be able to create a flowchart of mixtures using the vocabulary above.<br>• Be able to give examples of each. |
| Matter | Anything that has mass and takes up space (volume). Everything in the universe is made up of matter. |
| Atoms | The "building blocks" of matter. The basic particle from which all elements are made.<br>• Atoms are made of protons, neutrons, and electrons. |
| Elements | A pure substance that cannot be broken down into any other substance by chemical or physical means. Can "combine to make a chemical compound, or stand alone as they are."<br>• Be able to give examples of elements that "stand alone" (gold, neon, etc.) |
| Molecules | A group of two or more atoms that are held together by chemical bonds. |
| Compound | A pure substance made of two or more elements that are chemically combined. Water is an example of a compound made up of tiny water molecules. |

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| Reading the Periodic Table | Be able to identify the following from the periodic table:<br>• Atomic Number<br>• Atomic Mass<br>• Chemical Symbol |
| Planetary Drawings/Structure of an Atom | Be able to draw and label a planetary model of an atom. To be able to do this you need to remember the three rules:<br>1. The atomic number is equal to the number of protons.<br>2. The number of protons is equal to the number of electrons.<br>3. The atomic mass minus the atomic number is equal to the number of neutrons. |